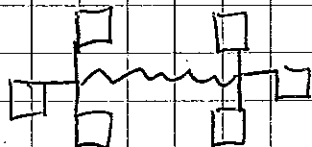


A, B, C with $n \geq 3$ C=C bonds at the functionalized ends can yield materials as follows using similar reaction conditions as those described on pg 14.

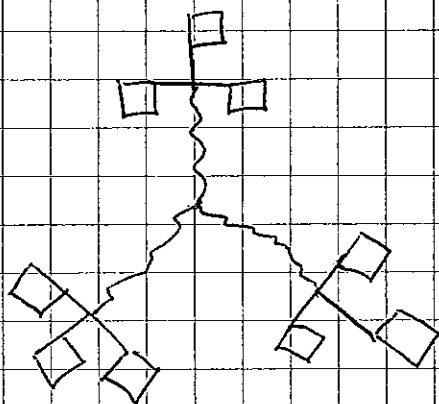
For "A" with $n \geq 3$ C=C bonds ($n=1$) as formed from either "1" or "2"



for "B" with 3 C=C bonds as formed from either "1" or "2" (3 on each functional end)



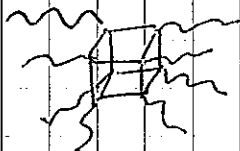
for "C" with 3 C=C bonds (on each functional end) as formed from either "1" or "2"



Similar chemistry using the hydrosilylation technique can be used to make PIB stars with a POSS core or crosslinked PIB networks where the crosslinking function is a POSS cage. This can be accomplished using the following POSS material in conjunction with A, B, or C.

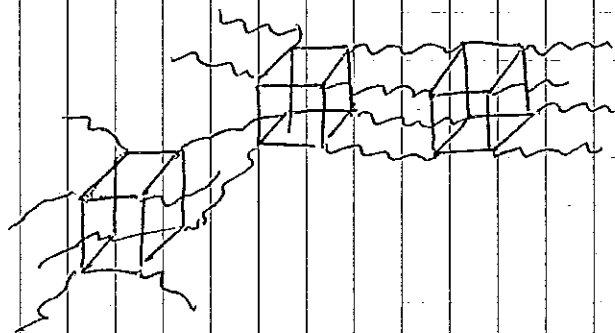
For example, an eight cornered PIB star core could be formed under conditions similar to those described on pg 14 with the use of the POSS material shown on pg 17 in place of that shown on pg 14.

For "A" with $R = CH_2-CH=CH_2$

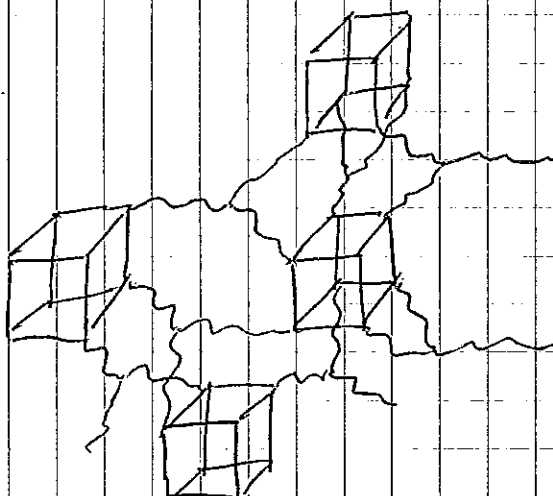


If "B" or "C" with $R = CH_2-CH=CH_2$ are used with the POSS described on pg 17. The following networks can be obtained

(for "B")

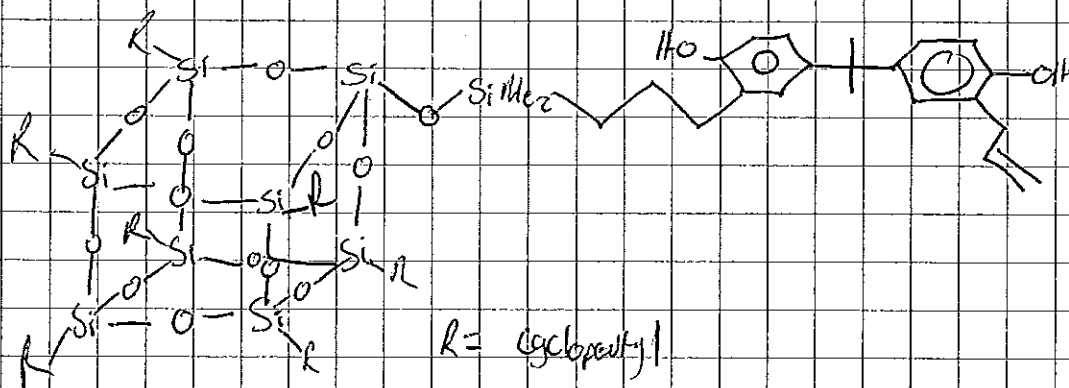


(for "C")



Important:

A POSS cage bearing two reactive Si-H groups might be formed by the reaction of two equivalents of $\text{H-Si}(\text{H}_2)_2\text{Cl}$ with one of the following pass materials



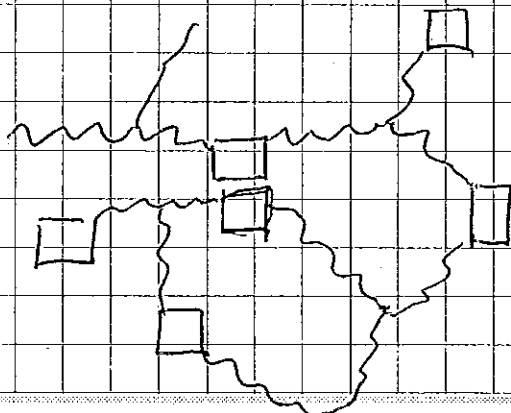
Reaction of such a POSS cage with "A" having $\text{R} = \text{CH}_2-\text{CH}=\text{CH}_2$ would lead to linear PEB with a POSS cage in the center of the polymer chain



Reaction of this same POSS cage with "B" having $\text{R} = \text{CH}_2-\text{CH}=\text{CH}_2$ would lead to linear PEB having several POSS cages in the polymer backbone at regular intervals

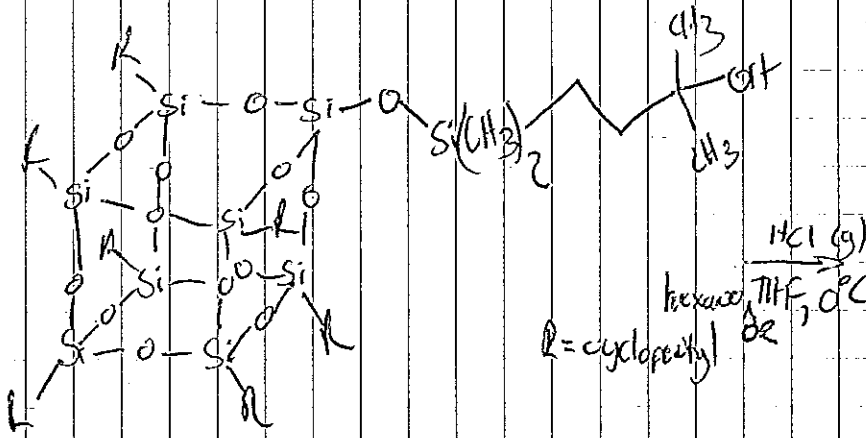


Reaction of this same POSS cage with "C" having $\text{R} = \text{CH}_2-\text{CH}=\text{CH}_2$ would lead to a network where each crosslink junction being a POSS cage



Important: Place card under blue copy.

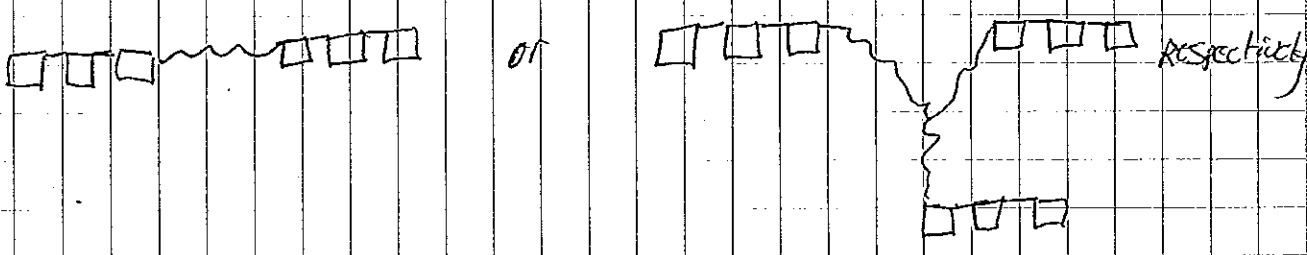
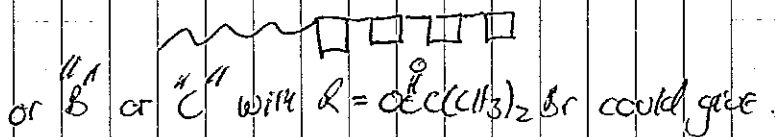
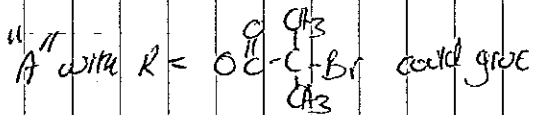
A POSS cage bearing a 3° (tertiary) alcohol could be converted to a 3° chloride and then be used as an initiator for PIB polymerization under living conditions. This might be accomplished to the following structure.



Finally, A, B, and C could bear other functional groups such as NH_2 that could undergo reaction with POSS cages bearing single or multiple epoxide groups to form materials similar to those previously described in analogous hydrosilylation reactions. There are many other possible living reactions including PIB-trialkoxysilyl / POSS-trialkoxysilyl, PIB-carboxylic acid / POSS-OH, PIB-OH / POSS-epoxide.

Furthermore if A, B, and C bear methacryloyl groups, blocks of POSS ~~and~~ bearing monomercaptoyl groups can be attached using ATRP.

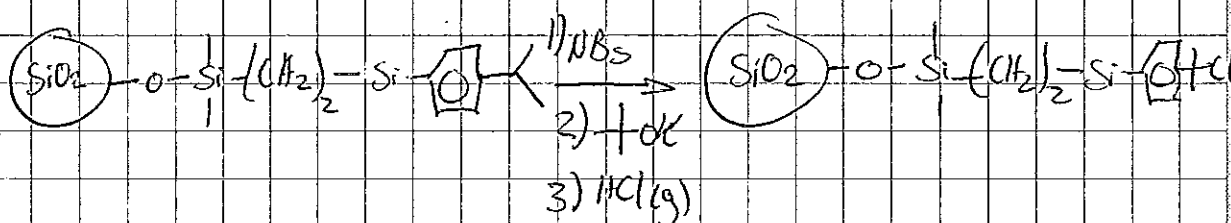
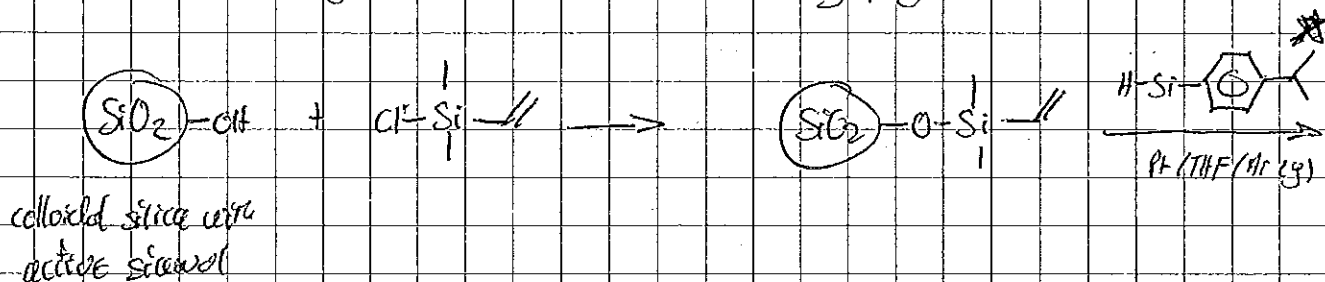
For example



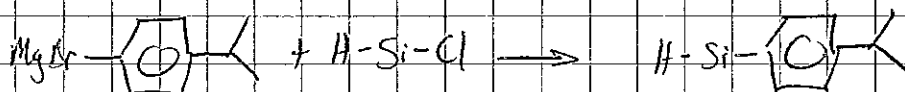
Important:

Sharda Sharma 04.20.00
 Laboratory Research

Although the direct functionalization of silica has been done prior in this group another method that I found indirect/possible functionalization could be used to functionalize colloidal silica with groups that can initiate the living polymerization of PB.



★



Similar chemistry could be used to functionalize colloidal silica with TMPCl.

Pias was synthesizing dicumyl chloride from commercially available dicumyl alcohol and HCl. The typical reaction conditions are as follows: $\text{C}_6\text{H}_5\text{Li}_2$ is added to a flask, next the required charge of HCl is added then with vigorous stirring (at room temp) the required amount of dicumyl alcohol is added. The reaction mixture is allowed to stir for several hours before the magnesium chloride layer is isolated (use separating funnel) and the product is isolated *vis a vis* distillation, however, Pias's product was yellow in color. He feared that this color was due to an impurity to the dicumyl alcohol. He attempted to remedy this by several recrystallizations of the dicumyl alcohol from ethanol. He conducted 3 such recrystallizations of the dicumyl alcohol before reacting it to form the chloride. Still, the product was slightly yellow. I decided to study the solubility of the alcohol and his crude dicumyl chloride in various common solvents.